

## General Principles and Processes of Isolation of Elements Class 12

### Chapter 6 General Principles and Processes of Isolation of Elements Exercise Solutions

In text : Solutions of Questions on Page Number : 150

Q1 :

Which of the ores mentioned in Table 6.1 can be concentrated by magnetic separation method?

Answer :

If the ore or the gangue can be attracted by the magnetic field, then the ore can be concentrated by the process of magnetic separation. Among the ores mentioned in table 6.1, the ores of iron such as haematite ( $\text{Fe}_2\text{O}_3$ ), magnetite ( $\text{Fe}_3\text{O}_4$ ), siderite ( $\text{FeCO}_3$ ), and iron pyrites ( $\text{FeS}_2$ ) can be separated by the process of magnetic separation.

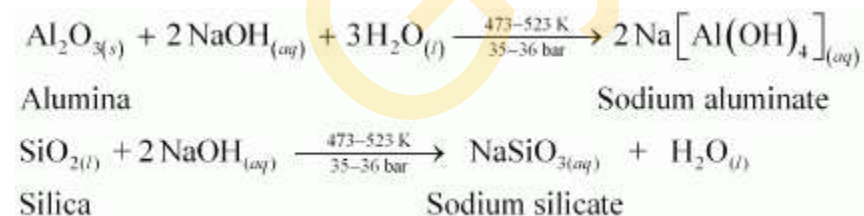
Q2 :

What is the significance of leaching in the extraction of aluminium?

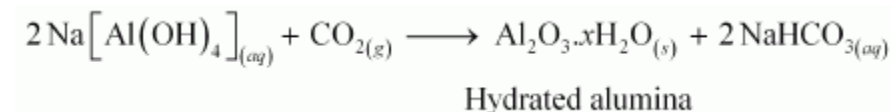
Answer :

In the extraction of aluminium, the significance of leaching is to concentrate pure alumina ( $\text{Al}_2\text{O}_3$ ) from bauxite ore.

Bauxite usually contains silica, iron oxide, and titanium oxide as impurities. In the process of leaching, alumina is concentrated by digesting the powdered ore with a concentrated solution of NaOH at 473-523 K and 35-36 bar. Under these conditions, alumina ( $\text{Al}_2\text{O}_3$ ) dissolves as sodium meta-aluminate and silica ( $\text{SiO}_2$ ) dissolves as sodium silicate leaving the impurities behind.



The impurities are then filtered and the solution is neutralized by passing  $\text{CO}_2$  gas. In this process, hydrated  $\text{Al}_2\text{O}_3$  gets precipitated and sodium silicate remains in the solution. Precipitation is induced by seeding the solution with freshly prepared samples of hydrated  $\text{Al}_2\text{O}_3$ .

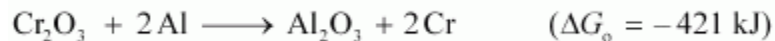


Hydrated alumina thus obtained is filtered, dried, and heated to give back pure alumina ( $\text{Al}_2\text{O}_3$ ).



**Q3 :**

**The reaction,**



**is thermodynamically feasible as is apparent from the Gibbs energy value.**

**Why does it not take place at room temperature?**

**Answer :**

The change in Gibbs energy is related to the equilibrium constant,  $K$  as

$$\Delta G = -RT \ln K$$

At room temperature, all reactants and products of the given reaction are in the solid state. As a result, equilibrium does not exist between the reactants and the products. Hence, the reaction does not take place at room temperature. However, at a higher temperature, chromium melts and the reaction takes place.

We also know that according to the equation

$$\Delta G = \Delta H - T\Delta S,$$

Increasing the temperature increases the value of  $T\Delta S$ , making the value of  $\Delta G$  more and more negative. Therefore, the reaction becomes more and more feasible as the temperature is increased.

**Q4 :**

**The reaction,**



**is thermodynamically feasible as is apparent from the Gibbs energy value.**

**Why does it not take place at room temperature?**

**Answer :**

The change in Gibbs energy is related to the equilibrium constant,  $K$  as

$$\Delta G = -RT \ln K$$

At room temperature, all reactants and products of the given reaction are in the solid state. As a result, equilibrium does not exist between the reactants and the products. Hence, the reaction does not take place at room temperature. However, at a higher temperature, chromium melts and the reaction takes place.

We also know that according to the equation

$$\Delta G = \Delta H - T\Delta S,$$

Increasing the temperature increases the value of  $T\Delta S$ , making the value of  $\Delta G$  more and more negative. Therefore, the reaction becomes more and more feasible as the temperature is increased.

<< Previous Chapter 5 : Surface Chemistry Next Chapter 7 : The p-Block Elements >>

**Exercise :** Solutions of Questions on Page Number : 163

**Q1 :**

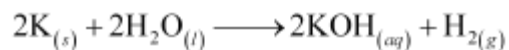
**Copper can be extracted by hydrometallurgy but not zinc. Explain.**

**Answer :**

The reduction potentials of zinc and iron are lower than that of copper. In hydrometallurgy, zinc and iron can be used to displace copper from their solution.



But to displace zinc, more reactive metals i.e., metals having lower reduction potentials than zinc such as Mg, Ca, K, etc. are required. But all these metals react with water with the evolution of  $\text{H}_2$  gas.



As a result, these metals cannot be used in hydrometallurgy to extract zinc.

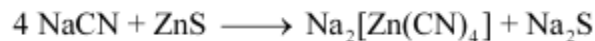
Hence, copper can be extracted by hydrometallurgy but not zinc.

**Q2 :**

**What is the role of depressant in froth floatation process?**

**Answer :**

In the froth floatation process, the role of the depressants is to separate two sulphide ores by selectively preventing one ore from forming froth. For example, to separate two sulphide ores ( $\text{ZnS}$  and  $\text{PbS}$ ),  $\text{NaCN}$  is used as a depressant which selectively allows  $\text{PbS}$  to come with froth, but prevents  $\text{ZnS}$  from coming to froth. This happens because  $\text{NaCN}$  reacts with  $\text{ZnS}$  to form  $\text{Na}_2[\text{Zn}(\text{CN})_4]$ .



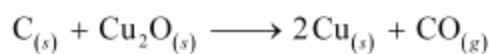
**Q3 :**

**Why is the extraction of copper from pyrites more difficult than that from its oxide ore through reduction?**

**Answer :**

The Gibbs free energy of formation ( $\Delta G$ ) of  $\text{Cu}_2\text{S}$  is less than that of  $\text{H}_2\text{S}$  and  $\text{CS}_2$ . Therefore,  $\text{H}_2$  and  $\text{C}$  cannot reduce  $\text{Cu}_2\text{S}$  to  $\text{Cu}$ .

On the other hand, the Gibbs free energy of formation of  $\text{Cu}_2\text{O}$  is greater than that of  $\text{CO}$ . Hence,  $\text{C}$  can reduce  $\text{Cu}_2\text{O}$  to  $\text{Cu}$ .



Hence, the extraction of copper from its pyrite ore is difficult than from its oxide ore through reduction.

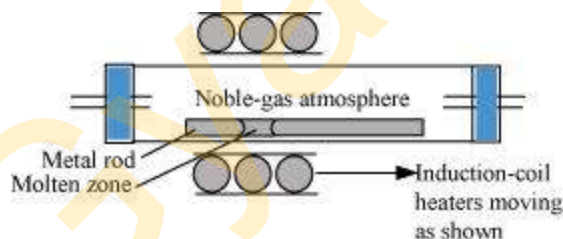
**Q4 :**

**Explain: (i) Zone refining (ii) Column chromatography.**

**Answer :**

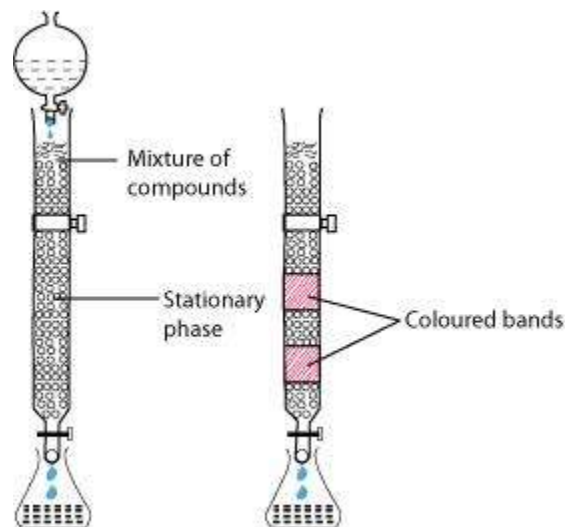
**(i) Zone refining:**

This method is based on the principle that impurities are more soluble in the molten state of metal (the melt) than in the solid state. In the process of zone refining, a circular mobile heater is fixed at one end of a rod of impure metal. As the heater moves, the molten zone of the rod also moves with it. As a result, pure metal crystallizes out of the melt and the impurities pass onto the adjacent molten zone. This process is repeated several times, which leads to the segregation of impurities at one end of the rod. Then, the end with the impurities is cut off. Silicon, boron, gallium, indium etc. can be purified by this process.



**(ii) Column chromatography:**

Column chromatography is a technique used to separate different components of a mixture. It is a very useful technique used for the purification of elements available in minute quantities. It is also used to remove the impurities that are not very different in chemical properties from the element to be purified. Chromatography is based on the principle that different components of a mixture are differently adsorbed on an adsorbent. In chromatography, there are two phases: mobile phase and stationary phase. The stationary phase is immobile and immiscible.  $\text{Al}_2\text{O}_3$  column is usually used as the stationary phase in column chromatography. The mobile phase may be a gas, liquid, or supercritical fluid in which the sample extract is dissolved. Then, the mobile phase is forced to move through the stationary phase. The component that is more strongly adsorbed on the column takes a longer time to travel through it than the component that is weakly adsorbed. The adsorbed components are then removed (eluted) using a suitable solvent (eluant).



**Q5 :**

**Out of C and CO, which is a better reducing agent at 673 K?**

**Answer :**

At 673 K, the value of  $\Delta G_{(CO,CO_2)}$  is less than that of  $\Delta G_{(C,CO)}$ . Therefore, CO can be oxidised more easily to  $CO_2$  than C to CO. Hence, CO is a better reducing agent than C at 673 K.

**Q6 :**

**Name the common elements present in the anode mud in electrolytic refining of copper. Why are they so present ?**

**Answer :**

In electrolytic refining of copper, the common elements present in anode mud are selenium, tellurium, silver, gold, platinum, and antimony.

These elements are very less reactive and are not affected during the purification process. Hence, they settle down below the anode as anode mud.

**Q7 :**

**Write down the reactions taking place in different zones in the blast furnace during the extraction of iron.**

**Answer :**

During the extraction of iron, the reduction of iron oxides takes place in the blast furnace. In this process, hot air is blown from the bottom of the furnace and coke is burnt to raise the temperature up to 2200 K in the lower portion itself. The temperature is lower in the upper part. Thus, it is the lower part where the reduction of iron oxides ( $\text{Fe}_2\text{O}_3$  and  $\text{Fe}_3\text{O}_4$ ) takes place.

The reactions taking place in the lower temperature range (500 - 800 K) in the blast furnace are:



**Q8 :**

**Write chemical reactions taking place in the extraction of zinc from zinc blende.**

**Answer :**

The different steps involved in the extraction of zinc from zinc blende ( $\text{ZnS}$ ) are given below:

**(i) Concentration of ore**

First, the gangue from zinc blende is removed by the froth floatation method.

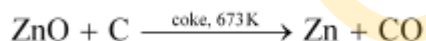
**(ii) Conversion to oxide (Roasting)**

Sulphide ore is converted into oxide by the process of roasting. In this process,  $\text{ZnS}$  is heated in a regular supply of air in a furnace at a temperature, which is below the melting point of  $\text{Zn}$ .



**(iii) Extraction of zinc from zinc oxide (Reduction)**

Zinc is extracted from zinc oxide by the process of reduction. The reduction of zinc oxide is carried out by mixing it with powdered coke and then, heating it at 673 K.



**(iv) Electrolytic Refining**

Zinc can be refined by the process of electrolytic refining. In this process, impure zinc is made the anode while a pure copper strip is made the cathode. The electrolyte used is an acidified solution of zinc sulphate ( $\text{ZnSO}_4$ ). Electrolysis results in the transfer of zinc in pure form from the anode to the cathode.

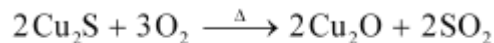
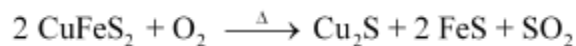


**Q9 :**

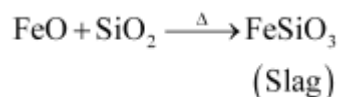
**State the role of silica in the metallurgy of copper.**

**Answer :**

During the roasting of pyrite ore, a mixture of FeO and Cu<sub>2</sub>O is obtained.



The role of silica in the metallurgy of copper is to remove the iron oxide obtained during the process of roasting as 'slag'. If the sulphide ore of copper contains iron, then silica (SiO<sub>2</sub>) is added as flux before roasting. Then, FeO combines with silica to form iron silicate, FeSiO<sub>3</sub> (slag).



**Q10 :**

**What is meant by the term "chromatography"?**

**Answer :**

Chromatography is a collective term used for a family of laboratory techniques for the separation of mixtures. The term is derived from Greek words 'chroma' meaning 'colour' and 'graphein' meaning 'to write'. Chromatographic techniques are based on the principle that different components are absorbed differently on an absorbent. There are several chromatographic techniques such as paper chromatography, column chromatography, gas chromatography, etc.

**Q11 :**

**What criterion is followed for the selection of the stationary phase in chromatography?**

**Answer :**

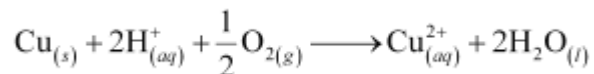
The stationary phase is selected in such a way that the components of the sample have different solubility's in the phase. Hence, different components have different rates of movement through the stationary phase and as a result, can be separated from each other.

**Q12 :**

**How is leaching carried out in case of low grade copper ores?**

**Answer :**

In case of low grade copper ores, leaching is carried out using acid or bacteria in the presence of air. In this process, copper goes into the solution as  $\text{Cu}^{2+}$  ions.



The resulting solution is treated with scrap iron or  $\text{H}_2$  to get metallic copper.



**Q13 :**

**Why is zinc not extracted from zinc oxide through reduction using CO?**

**Answer :**

The standard Gibbs free energy of formation of ZnO from Zn

is lower than that of  $\text{CO}_2$  from CO. Therefore, CO cannot reduce ZnO to Zn. Hence, Zn is not extracted from ZnO through reduction using CO.

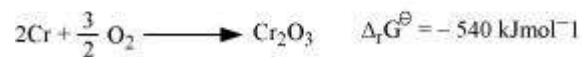
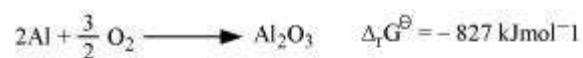
**Q14 :**

The value of  $\Delta_f G^\ominus$  for formation of  $\text{Cr}_2\text{O}_3$  is  $-540 \text{ kJmol}^{-1}$  and that of  $\text{Al}_2\text{O}_3$  is  $-827 \text{ kJmol}^{-1}$ . Is the reduction of  $\text{Cr}_2\text{O}_3$  possible with Al?

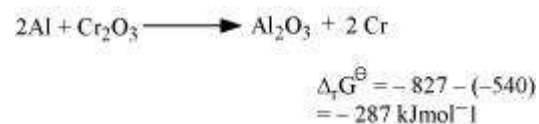
**Answer :**

The value of  $\Delta_f G^\ominus$  for the formation of  $\text{Cr}_2\text{O}_3$  from Cr ( $-540 \text{ kJmol}^{-1}$ ) is higher than that of  $\text{Al}_2\text{O}_3$  from Al ( $-827 \text{ kJmol}^{-1}$ ). Therefore, Al can reduce  $\text{Cr}_2\text{O}_3$  to Cr. Hence, the reduction of  $\text{Cr}_2\text{O}_3$  with Al is possible.

Alternatively,



Subtracting equation (ii) from (i), we have



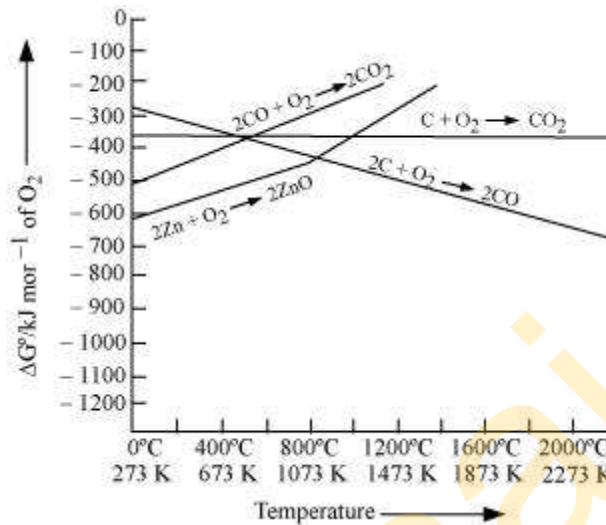
As  $\Delta_r G^\ominus$  for the reduction reaction of  $\text{Cr}_2\text{O}_3$  by Al is negative, this reaction is possible.



Q15 :

Out of C and CO, which is a better reducing agent for ZnO ?

Answer :



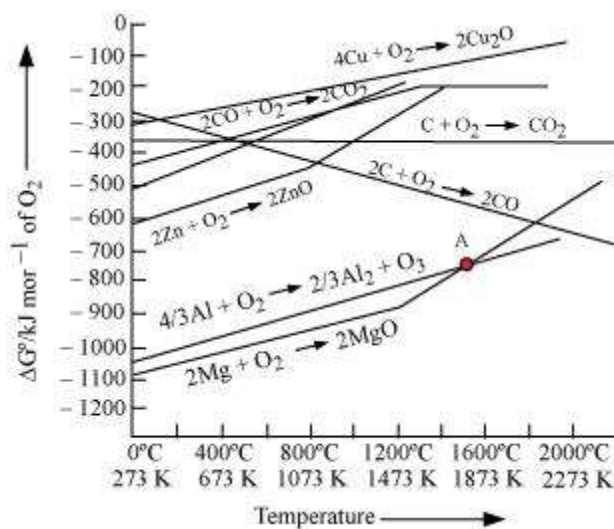
Reduction of ZnO to Zn is usually carried out at 1673 K. From the above figure, it can be observed that above 1073 K, the Gibbs free energy of formation of CO from C and above 1273 K, the Gibbs free energy of formation of  $\text{CO}_2$  from C is lesser than the Gibbs free energy of formation of ZnO. Therefore, C can easily reduce ZnO to Zn.

On the other hand, the Gibbs free energy of formation of  $\text{CO}_2$  from CO is always higher than the Gibbs free energy of formation of ZnO. Therefore, CO cannot reduce ZnO. Hence, C is a better reducing agent than CO for reducing ZnO.

Q16 :

The choice of a reducing agent in a particular case depends on thermodynamic factor. How far do you agree with this statement? Support your opinion with two examples.

Answer :



The above figure is a plot of Gibbs energy ( $\Delta G^\ominus$ ) vs. T for formation of some oxides.

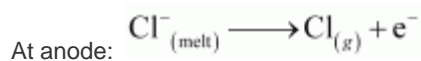
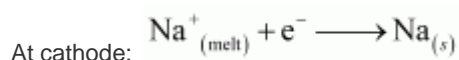
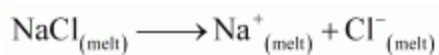
It can be observed from the above graph that a metal can reduce the oxide of other metals, if the standard free energy of formation ( $\Delta_f G^\ominus$ ) of the oxide of the former is more negative than the latter. For example, since  $\Delta_f G^\ominus(\text{Al}, \text{Al}_2\text{O}_3)$  is more negative than  $\Delta_f G^\ominus(\text{Cu}, \text{Cu}_2\text{O})$ , Al can reduce  $\text{Cu}_2\text{O}$  to Cu, but Cu cannot reduce  $\text{Al}_2\text{O}_3$ . Similarly, Mg can reduce ZnO to Zn, but Zn cannot reduce MgO because  $\Delta_f G^\ominus(\text{Mg}, \text{MgO})$  is more negative than  $\Delta_f G^\ominus(\text{Zn}, \text{ZnO})$ .

**Q17 :**

**Name the processes from which chlorine is obtained as a by-product. What will happen if an aqueous solution of NaCl is subjected to electrolysis?**

**Answer :**

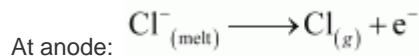
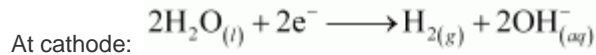
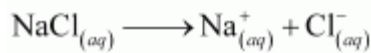
In the electrolysis of molten NaCl,  $\text{Cl}_2$  is obtained at the anode as a by product.



The overall reaction is as follows:



If an aqueous solution of NaCl is electrolyzed, Cl<sub>2</sub> will be obtained at the anode but at the cathode, H<sub>2</sub> will be obtained (instead of Na). This is because the standard reduction potential of Na (E° = - 2.71 V) is more negative than that of H<sub>2</sub>O (E° = - 0.83 V). Hence, H<sub>2</sub>O will get preference to get reduced at the cathode and as a result, H<sub>2</sub> is evolved.

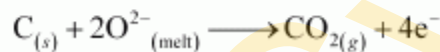
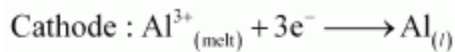


**Q18 :**

**What is the role of graphite rod in the electrometallurgy of aluminium?**

**Answer :**

In the electrometallurgy of aluminium, a fused mixture of purified alumina (Al<sub>2</sub>O<sub>3</sub>), cryolite (Na<sub>3</sub>AlF<sub>6</sub>) and fluorspar (CaF<sub>2</sub>) is electrolysed. In this electrolysis, graphite is used as the anode and graphite-lined iron is used as the cathode. During the electrolysis, Al is liberated at the cathode, while CO and CO<sub>2</sub> are liberated at the anode, according to the following equation.



If a metal is used instead of graphite as the anode, then O<sub>2</sub> will be liberated. This will not only oxidise the metal of the electrode, but also convert some of the Al liberated at the cathode back into Al<sub>2</sub>O<sub>3</sub>. Hence, graphite is used for preventing the formation of O<sub>2</sub> at the anode. Moreover, graphite is cheaper than other metals.

**Q19 :**

**Outline the principles of refining of metals by the following methods:**

**(i) Zone refining**

**(ii) Electrolytic refining**

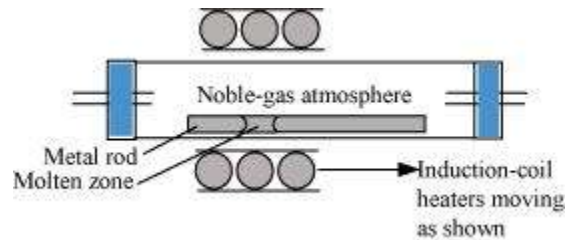
**(iii) Vapour phase refining**

**Answer :**

**(i) Zone refining:**

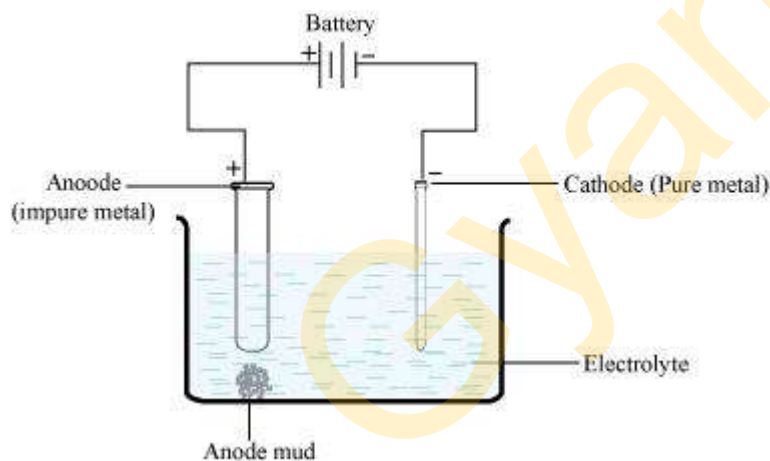
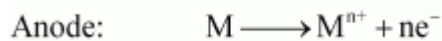
This method is based on the principle that impurities are more soluble in the molten state of metal (the melt) than in the solid state. In the process of zone refining, a circular mobile heater is fixed at one end of a rod of impure metal.

As the heater moves, the molten zone of the rod also moves along with it. As a result, pure metal crystallizes out of the melt and the impurities pass to the adjacent molten zone. This process is repeated several times, which leads to the segregation of impurities at one end of the rod. Then, the end with the impurities is cut off. Silicon, boron, gallium, indium etc. can be purified by this process.



### (ii) Electrolytic refining;

Electrolytic refining is the process of refining impure metals by using electricity. In this process, impure metal is made the anode and a strip of pure metal is made the cathode. A solution of a soluble salt of the same metal is taken as the electrolyte. When an electric current is passed, metal ions from the electrolyte are deposited at the cathode as pure metal and the impure metal from the anode dissolves into the electrolyte in the form of ions. The impurities present in the impure metal gets collected below the anode. This is known as anode mud.



### (iii) Vapour phase refining

Vapour phase refining is the process of refining metal by converting it into its volatile compound and then, decomposing it to obtain a pure metal. To carry out this process,

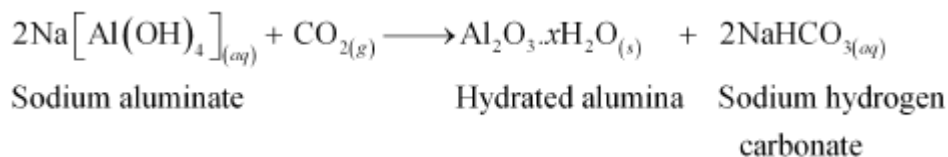
- (i) the metal should form a volatile compound with an available reagent, and
- (ii) the volatile compound should be easily decomposable so that the metal can be easily recovered.

Nickel, zirconium, and titanium are refined using this method.

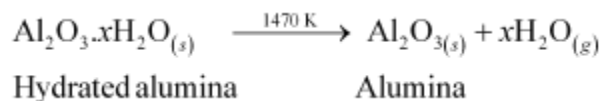
Q20 :

Predict conditions under which Al might be expected to reduce MgO.





During this process, sodium silicate remains in the solution. The obtained hydrated alumina is filtered, dried, and heated to get back pure alumina.

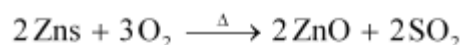


**Q23 :**

**Giving examples, differentiate between 'roasting' and 'calcination'.**

**Answer :**

Roasting is the process of converting sulphide ores to oxides by heating the ores in a regular supply of air at a temperature below the melting point of the metal. For example, sulphide ores of Zn, Pb, and Cu are converted to their respective oxides by this process.



Zinc blende

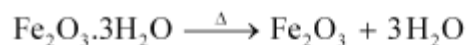


Galena

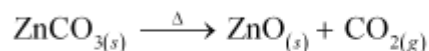


Copper glance

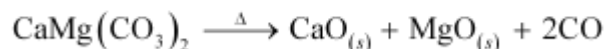
On the other hand, calcination is the process of converting hydroxide and carbonate ores to oxides by heating the ores either in the absence or in a limited supply of air at a temperature below the melting point of the metal. This process causes the escaping of volatile matter leaving behind the metal oxide. For example, hydroxide of Fe, carbonates of Zn, Ca, Mg are converted to their respective oxides by this process.



Limonite



Calamine



Dolomite

**Q24 :**

**How is 'cast iron' different from 'pig iron'?**

**Answer :**

The iron obtained from blast furnaces is known as pig iron. It contains around 4% carbon and many impurities such as S, P, Si, Mn in smaller amounts.

Cast iron is obtained by melting pig iron and coke using a hot air blast. It contains a lower amount of carbon (3%) than pig iron. Unlike pig iron, cast iron is extremely hard and brittle.

**Q25 :**

**Differentiate between "minerals" and "ores".**

**Answer :**

Minerals are naturally occurring chemical substances containing metals. They are found in the Earth's crust and are obtained by mining.

Ores are rocks and minerals viable to be used as a source of metal.

For example, there are many minerals containing zinc, but zinc cannot be extracted profitably (conveniently and economically) from all these minerals.

Zinc can be obtained from zinc blende (ZnS), calamine (ZnCO<sub>3</sub>), Zincite (ZnO) etc.

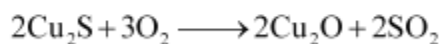
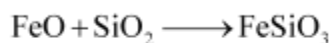
Thus, these minerals are called ores of zinc.

**Q26 :**

**Why copper *matte* is put in silica lined converter?**

**Answer :**

Copper *matte* contains Cu<sub>2</sub>S and FeS. Copper *matte* is put in a silica-lined converter to remove the remaining FeO and FeS present in the *matte* as slag (FeSiO<sub>3</sub>). Also, some silica is added to the silica-lined converter. Then, a hot air blast is blown. As a result, the remaining FeS and FeO are converted to iron silicate (FeSiO<sub>3</sub>) and Cu<sub>2</sub>S is converted into metallic copper.



**Q27 :**

**What is the role of cryolite in the metallurgy of aluminium?**

**Answer :**

Cryolite ( $\text{Na}_3\text{AlF}_6$ ) has two roles in the metallurgy of aluminium:

1. To decrease the melting point of the mixture from 2323 K to 1140 K.
2. To increase the electrical conductivity of  $\text{Al}_2\text{O}_3$ .

Gyanajal